

Table F

1. When $\text{PbI}_2(\text{s})$ is added to $\text{Na}_2\text{CO}_3(\text{aq})$, a white precipitate is formed. According to Reference Table F, the white precipitate most likely is
- A) KNO_3 B) PbCO_3
C) NaI D) Na_2CO_3
2. Which compound is insoluble in water?
- A) calcium bromide
B) potassium bromide
C) silver bromide
D) sodium bromide
3. Which barium salt is insoluble in water?
- A) BaCO_3 B) BaCl_2
C) $\text{Ba}(\text{ClO}_4)_2$ D) $\text{Ba}(\text{NO}_3)_2$
4. Which compound is insoluble in water?
- A) BaSO_4 B) CaCrO_4
C) KClO_3 D) Na_2S
5. Which ion, when combined with chloride ions, Cl^- , forms an insoluble substance in water?
- A) Fe^{2+} B) Mg^{2+} C) Pb^{2+} D) Zn^{2+}
6. Based on Reference Table F, which of these saturated solutions has the lowest concentration of dissolved ions?
- A) $\text{NaCl}(\text{aq})$ B) $\text{MgCl}_2(\text{aq})$
C) $\text{NiCl}_2(\text{aq})$ D) $\text{AgCl}(\text{aq})$
7. According to Reference Table F, which of these compounds is most soluble at 298 K and 1 atm?
- A) AgNO_3 B) AgCl
C) PbCrO_4 D) PbCO_3
8. According to Reference Table F, which substance is most soluble?
- A) AgI B) CaSO_4
C) PbCl_2 D) $(\text{NH}_4)_2\text{CO}_3$
9. Based on Reference Table F, which of the following saturated solutions would be the least concentrated?
- A) sodium sulfate
B) potassium sulfate
C) copper (II) sulfate
D) barium sulfate
10. Based on Reference Table F, which salt is least soluble?
- A) FeCO_3 B) Na_2CO_3
C) BaCl_2 D) CaCl_2